

Research Article

Gel Point Determination of TEOS-Based Polymeric Materials with Application on Conservation of Cultural Heritage Buildings

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Here we present measurements on the gel time of inorganic-organic materials used for stone preservation by means of rheology and dynamic light scattering. Our hybrid material is composed of tetraethyl orthosilicate (TEOS) and polydimethylsiloxane (PDMS) using a nonionic surfactant (n-octylamine) as a template. Moreover, zinc oxide (ZnO) nanoparticles are dispersed in the medium with the aim of obtaining a nanocomposite with potential biocide properties. In our case, we use the ZnO particles as tracers to infer from their scattered intensity mechanical information of the suspending medium. We have found that dynamic light scattering experiments provide similar information on the gelling time, about 30 hours, to that obtained from rotational rheology and oscillatory rheology. This result confirms the validity of light scattering, which is a noninvasive technique, to characterize mechanical properties of time evolving hybrid materials through nonperturbative and well-controlled experiments.

1. Introduction

Stone coating has been extensively used lastly to conserve our artistic stonework. In this direction, an enormous effort has been done to design materials that, without changing the visual appearance of the stone, provide it with consolidant properties and protection against environmental hazards such as erosion, humidity, and microbial attack [1]. In this regard, organically modified silicates (ORMOSILs) seem to fulfill most of these requirements, except antimicrobial activity, and they have been successfully used for stone preservation [2, 3]. As with all alkoxy silane-based products, these materials undergo a sol-gel transition. They are applied on the surface when the material is a liquid with low viscosity, and then the product polymerizes spontaneously inside the porous structure of the stone creating a polymer network of silica [4]. During polymerization two reactions take place: hydrolysis of alkoxy groups to generate silanol ones and condensation between silanol groups of the product

and/or between these groups and those present in the silicate minerals. TEOS polymeric precursor is the most commonly used for stone consolidation, mixed with solvents to homogenize the water-alkoxy silane mixture, and with a catalyst to increase the rate of the hydrolysis reaction. In our case, the hybridization of the silica gel is done with polydimethylsiloxane (PDMS) since it increases flexibility and strength of the resulting coating [5, 6] and accelerates the gelation process [7, 8]. Moreover, adding PDMS to the silicon alkoxide does not increase the viscosity of the initial sol, which makes its application over the stone and its penetration equally facile. However, most of the gels obtained from commercial formulations are highly brittle and crack easily. This inconvenience is circumvented by adding a nonionic surfactant (n-octylamine) to the synthesis as shown by the group led by Mosquera [9, 10]. The presence of this nonionic surfactant accelerates the gelling process [8] and avoids gel fracture during the drying stage due to the formation of a more uniform size distribution of mesopores

[11]. In this way, ORMOSIL products with hydrofugant and consolidant properties can be obtained by mixing TEOS and a poly(dimethylsiloxane) with terminal OH groups, PDMS, under the presence of a surfactant [10, 12]. To provide the material with antimicrobial activity, here we propose the addition of ZnO nanoparticles [13, 14] to the aforementioned polymeric formulation [12]. We select these nanoparticles due to their confirmed chemical stability, biocompatibility, and bioactivity [15] and to their limited use in materials for stone conservation. Moreover, these nanoparticles may also provide superhydrophobicity to the material [16], which is also highly desired for coatings used for stone protection. Recently, it has been shown that a similar product (prehydrolyzed TEOS, water, and n-octylamine) acquires biocide action after adding to the synthesis CuO nanoparticles [17].

The study of the sol-gel process existing in this kind of products has an enormous importance, in particular, the sol-gel transition point, to characterize the formation of the network structure that is responsible for the potential applications of such materials. Furthermore, the gelling time is a crucial parameter to evaluate their suitability since gelation cannot be too fast, to facilitate its application over the stone and its penetrability, nor too slow, to avoid evaporation of volatile molecules. During the sol-gel process, there is an enhancement of the viscosity due to the cluster formation, and then, at the gel point, the emergence of a three-dimensional network provokes a divergence of the viscosity and the appearance of elasticity in the material, process that is also known as “amorphous solidification” or “gelation” transition [18]. Indeed, after this point, it is said that a gel is formed, and an infinite cluster spans the sample that is capable of sustaining stress provoking a change of its macroscopic properties, where material elasticity overpasses viscosity. The determination of the gel point can be done employing several techniques. For example, Raman spectroscopy may provide useful information about the polymerization process underlying the gel formation, and it has been applied to several materials with coatings applications as polyalkyl siloxanes [19], biopolymers [20], among others. Alternatively, other techniques as rotational [21] or oscillatory [22, 23] rheological tests are required to follow the mechanical properties during gelation. However, rheology possesses some uncertainty in the determination of the gel point due to, on the one hand, the intrinsic difficulty of evaluating the divergence of one physical magnitude and, on the other hand, the shear applied to the material in this kind of measurements can perturb the sol-gel transition we want to characterize. An alternative technique to study the mechanical properties of a given medium is microrheology [24]. This technique is based on the fact that the motion of particles embedded in a given medium is related to its mechanical properties. In general, there are several techniques to obtain information on the dynamics of probe particles immersed in viscoelastic systems. One of them is based on dynamic light scattering [25]. Sometimes, the scatterers are introduced intentionally into the sample and, other times, the system possesses intrinsic inclusions, which can be used as probe particles. The advantage of microrheology is that the viscoelastic response of the medium is measured in the linear regime since the stress applied to

the material by the probe particle is due to its Brownian motion.

The previously mentioned techniques can be used to characterize the sol-gel process under interest, but they not always agree with each other [26]. To discard possible differences among different techniques, we characterize the gel point of our novel material by rheological measurements and light scattering techniques. The advantage of using the latter technique is that it is noninvasive and, in our case, we use the ZnO particles present in the system as tracer particles with no need of introducing additional optical probes. Our product could be a novel alternative since the sol-gel route here involved produces homogenous crack-free and transparent nanomaterials that could combine multiple protection mechanisms for building stone (consolidant, hydrofugant, and biocidal effects).

2. Materials and Methods

2.1. Sol-Gel Preparation. In this work, we synthesize silica-ZnO organic-inorganic nanocomposites by mixing ZnO nanoparticles with a silica oligomer in the presence of a non-ionic surfactant (n-octylamine) and of a hydroxyl-terminated polydimethylsiloxane. These nanocomposites integrate ZnO nanoparticles in a mesoporous silica-based matrix prepared according to a synthesis route previously published [12]. The nanomaterial starts from a sol containing TES 40 WN (from Wacker), which is a prehydrolyzed tetraethyl orthosilicate, in the presence of n-octylamine (from Aldrich). According to its technical data sheet, TES 40 WN (hereafter TES40) is a mixture of monomeric and oligomeric ethoxysilanes, whose average chain length is approximately 5 Si–O units. The ZnO nanopowders (from Nanostructured & Amorphous Materials, Inc.) have an average primary particle size of 20 nm provided by the supplier. PDMS (from Gelest) has a polymerization degree of 12 (molar mass 400–700) and an OH percentage ranging from 4.5 to 7.5% w/w. The synthesis of our organic-inorganic nanocomposite is as follows. Firstly, we prepare by vigorous stirring an aqueous solution of n-octylamine at a concentration of $1.57 \text{ mol}\cdot\text{dm}^{-3}$, which is significantly higher than its critical micellar concentration. Then the aqueous solution of n-octylamine is mixed with TES40 under high-power ultrasonic agitation ($60 \text{ W}\cdot\text{cm}^{-3}$) for 10 minutes. The resulting molar ratio of TES40 to n-octylamine and water are $1: 5.3 \times 10^{-4}$ and $1: 133.6 \times 10^{-4}$, respectively. Next, under this ultrasonic agitation, ZnO particles are incorporated up to a weight fraction of 0.005 % w/w. Finally, keeping the ultrasonic agitation, PDMS is added to the starting sol drop by drop up to a concentration of 10.4 % v/v.

2.2. Rheology Measurements. The viscosity and viscoelastic moduli of our suspensions are measured with an MCR 502 rheometer (Anton Paar) by using a cone-plate geometry of 50 mm of diameter and a cone angle of one degree. The steady flow curves to determine the shear viscosity are obtained by rotational tests with controlled shear rate. The time needed for the sample to reach constant conditions for each shear rate, and consequently to perform the measurement, is

determined automatically by the rheometer after recording constant consecutive values of the viscosity. The frequency dependence of the viscoelastic moduli is measured in the oscillatory mode in the linear viscoelastic regime by applying a strain of 0.3 %. This strain assures a linear response of the sample for all measurements, confirmed by the dependence of the storage and loss moduli with deformation measured at a frequency of 10 rad/s.

2.3. Light Scattering Technique. The size of the ZnO colloidal particles is measured with a Zetasizer ZS90 (Malvern Inst.). This device estimates the size of particles immersed in a fluid from the autocorrelation function of the scattered intensity by means of the well-known technique dynamic light scattering (DLS). The fluctuation of the intensity scattered by the particles is related to their motion and the intensity correlation function minus one, $g_I(\tau) - 1$, in case of colloidal systems composed of particles having the same size and immersed in purely viscous fluids, decays exponentially with time [27]:

$$g_I(\tau) - 1 = Ae^{-\tau/\tau_c} \quad (1)$$

where τ_c is the characteristic decay time of the exponential function and is equal to

$$\tau_c = \frac{1}{2D_0q^2} \quad (2)$$

and here q is the scattering vector given by

$$q = \frac{4\pi n}{\lambda} \sin\left(\frac{\theta}{2}\right). \quad (3)$$

The scattering vector depends on the refractive index, n , of the media where the particles are immersed, the wavelength of the laser, $\lambda = 632.8$ nm, and the scattering angle, $\theta = 90^\circ$. D_0 is the free particle diffusion coefficient given by the Stokes-Einstein relation that, in case of a sphere, takes the form

$$D_0 = \frac{\kappa_B T}{6\pi\eta a} \quad (4)$$

where T is the temperature of the suspending medium, η is its viscosity, and a is the particle radius. Note that this expression can be equally used to determine the size of probe particles, when the viscosity of the fluid is known, or to determine the viscosity of a certain fluid by using particles of known diameter. The latter application is the most straightforward case of the technique called microrheology.

3. Results and Discussion

First, we measure the size of ZnO nanoparticles at 25°C when they are immersed in TES40 and in the system of interest, that is, after adding PDMS and the n-octylamine solution to the TES40, immediately after preparation. In both cases, the estimated sizes and polydispersity index, IPD, are equal to 360 ± 30 nm and 0.20 ± 0.05 , respectively. From this result, we prove that the particles are forming nanoclusters of similar size that are uniformly distributed in the precursor media,

TABLE 1: Viscosities measured at T=25°C by rotational rheology of the main component of the sample, TES40, and after adding the different components, TES40+PDMS, TES40+PDMS+n-octylamine solution and TES40+PDMS+n-octylamine solution+ZnO particles.

Sample	η [mPa·s]
TES40	4.28±0.13
TES40+PDMS	4.44±0.07
TES40+PDMS+n-octylamine solution	4.56±0.14
TES40+PDMS+n-octylamine solution+ZnO	4.6±0.4

TES40, and that the addition of the rest of components to the product does not affect such homogeneity.

The size of the ZnO particles is determined by using the Stokes-Einstein relation, where the viscosity is needed. The viscosities of the different suspending media are measured by rotational rheology. We measure the viscosity of the main component of the media, TES40, and then we characterize how it changes as the rest of the reagents, PDMS, solution of n-octylamine, and colloidal particles are added. The mean viscosities measured at 25°C are shown in Table 1, where the error is estimated as the sample standard deviation. As we can see, the addition of the different components to the precursor media slightly augments its viscosity, keeping the material penetrability into the stone.

After the sample preparation, once the system is ultrasonicated, its mechanical properties are followed in time at a temperature of 25°C. By performing rotational tests, we observe that the viscosity does not depend on shear rate during the initial hours after preparation in the interval of shear rates analyzed, showing the typical behavior of Newtonian fluids. However, after around 28 hours, viscosity starts to show a slight dependence on shear rate due to the polymerization process. During these first 28 hours, the viscosity only augments slightly but, after this waiting time, the viscosity increases very abruptly as shown in Figure 1(b), where we plot the dependence on reaction time of the viscosity normalized by its initial value measured at zero time, just after sample preparation. This sudden increase of the viscosity on time is a hallmark of the sol-gel transition, caused by the formation of an interconnected network that spans the sample. In fact, it is by performing a power law fit to the data (continuous line in Figure 1(b)) by using the model

$$\frac{\eta(t)}{\eta_0} = \left(1 - \frac{t}{t_{gel}}\right)^{-s}. \quad (5)$$

We get that the gelation time $t_{gel} = (30.2 \pm 0.1)$ hours and $s = (0.72 \pm 0.02)$. This expression has been previously used to describe very well the time dependence of the viscosity in sol-gel transitions [28]. The value obtained for the critical exponent is quite close to the experimental value 0.75 found in silica-siloxane mixtures [29] and to the expected value of 0.7 predicted by the percolation theory [28].

Up to now, our characterization is based on rotational rheology where shear is applied to the sample and, consequently, this perturbation may alter the gelification process. For this reason, we also characterize the sol-gel transition

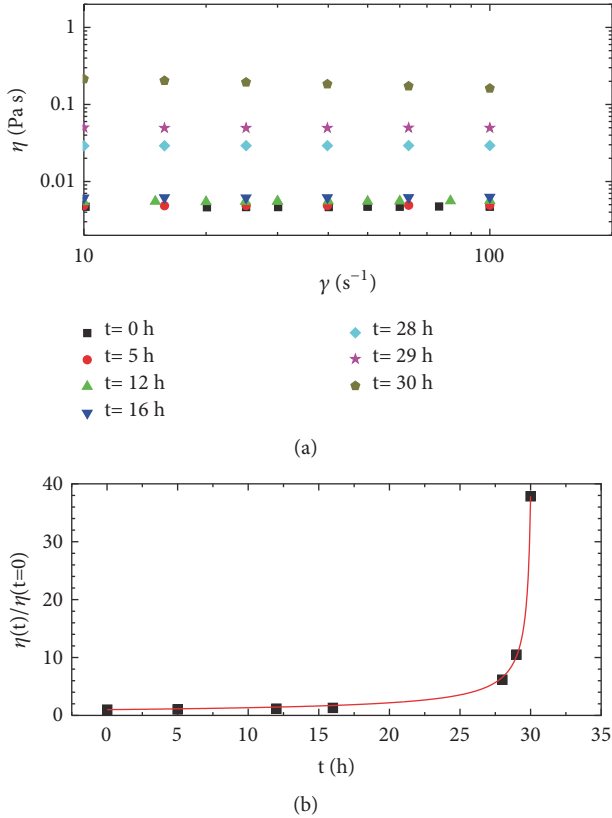


FIGURE 1: (a) Viscosity as a function of shear rate measured at different times after sample preparation at $T=25^\circ\text{C}$: black squares ($t = 0$ h), red circles ($t = 5$ h), green up triangles ($t = 12$ h), blue down triangles ($t = 16$ h), cyan diamonds ($t = 28$ h), pink stars ($t = 29$ h), and dark yellow pentagon ($t = 30$ h). (b) Normalized viscosity measured as a function of time (squares) and power law fit, $(1 - (t/t_{\text{gel}}))^{-s}$, to the experimental data, where t_{gel} and s are fitting parameters (continuous line).

by oscillatory rheology where we determine the frequency dependence of the viscoelastic moduli on the reaction time by applying an oscillatory stress whose amplitude is very small to assure the linear viscoelastic regime. The measured viscoelastic moduli are plotted as a function of frequency for different reaction times in Figure 2(a). For all the times investigated both moduli increase with frequency. At the initial times, the viscous modulus (hollow symbols) is higher than the elastic one (solid symbols) for all the frequencies under investigation up to a certain time, where the elastic component overcomes the viscous one, which shows the existence of a gel. This inversion point is placed between 31 and 32 hours and, precisely in this interval, we can observe in Figure 2(b) the power law dependence of the elastic and viscous moduli with frequency $G' \sim G'' \sim \omega^n$, where n is known as the critical viscoelastic exponent. This power law frequency dependence of the dynamic moduli resembles the behavior of a weak gel and serves to identify the formation of an incipient gel at the gel point [22, 23], indicating the presence of a three-dimensional network and, consequently, the transition from a liquid to a solid-like behavior.

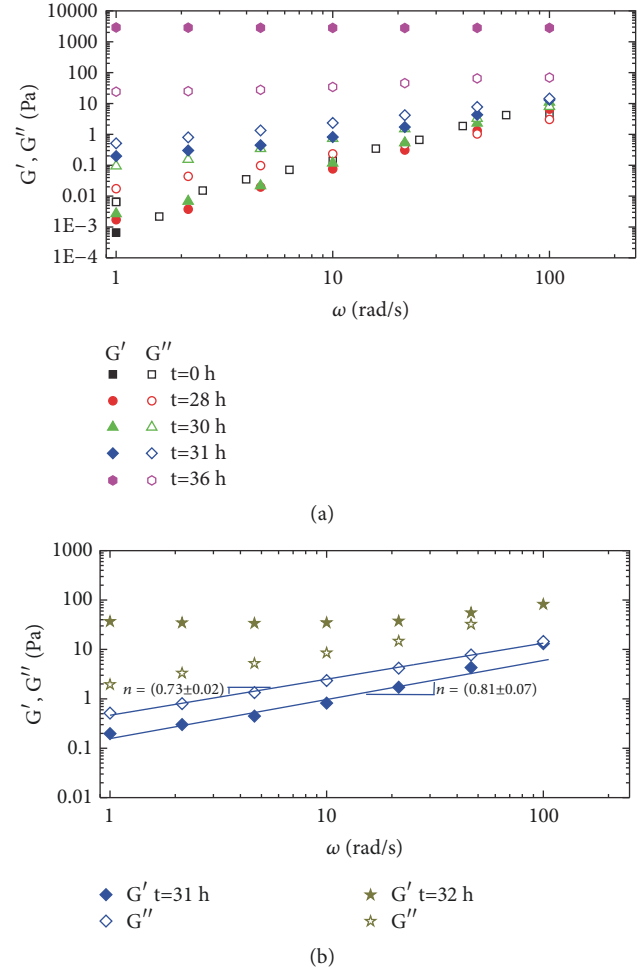


FIGURE 2: (a) Elastic (solid symbols) and viscous (hollow symbols) moduli measured as a function of frequency at different reaction times at $T=25^\circ\text{C}$: immediately after sample preparation (black squares), after 28 hours (red circles), after 30 hours (green triangles), and after 36 hours (pink hexagons). (b) Elastic and viscous moduli measured 31 hours after sample preparation (blue diamonds) along their corresponding fits to a power law $G \sim \omega^n$ (continuous lines) and 32 hours after sample preparation (dark yellow stars).

In our case, the critical viscoelastic exponents $n = (0.81 \pm 0.07)$ obtained from the frequency dependence of the elastic modulus, i.e., $G'(\omega) \sim \omega^{0.81 \pm 0.07}$ and $n = (0.73 \pm 0.02)$, from the frequency analysis of the viscous modulus, $G''(\omega) \sim \omega^{0.73 \pm 0.02}$, are in agreement with the interval of values found in the literature. For branched polymer, as in our case, it is expected that this value is in the interval that goes from 0.66 to 1 [30] and it is close to the value 0.67 derived from percolation theory [31]. The values so obtained for the critical exponents, considering their uncertainty, are consistent between them, which guarantees that the gel point is quite close to this time. In fact, the next measurement, performed after 32 hours, shows a solid-like behavior since for all frequencies investigated the elastic modulus is higher than the viscous one, and at low frequencies the elastic modulus tends to a finite value, indicative of the gel nature of the sample.

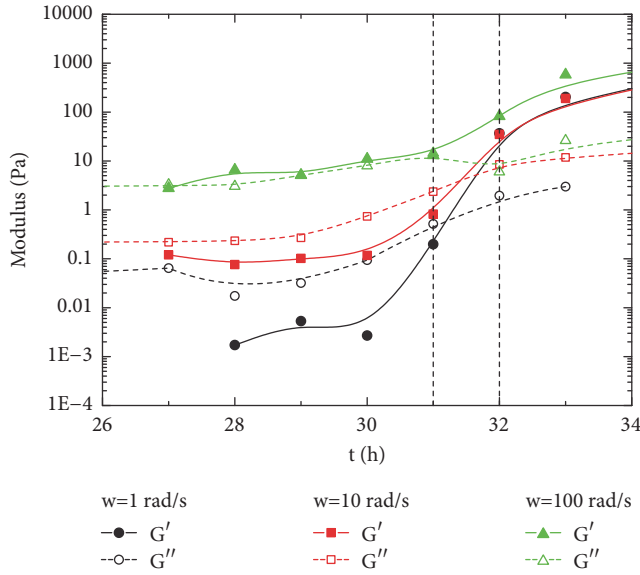


FIGURE 3: Elastic (solid symbols) and viscous (hollow symbols) moduli measured at frequencies of 1 rad/s (black circles), 10 rad/s (red squares), and 100 rad/s (green triangles) as a function of time. Lines are guides to the eye. Dashed vertical lines indicate the time interval where the crossover of $G' = G''$ occurs.

As the gel evolves, elasticity increases faster than viscosity at all frequencies and, after 36 hours, the system resembles the mechanical response of a fully developed network where the elastic modulus hardly depends on frequency and is two orders of magnitude larger than the viscous one, which are typical features of strong gels. From this finding, we can guarantee that the sol to gel transition occurs between 31 and 32 hours after sample preparation. This gel time also coincides with the time for which the elastic modulus (solid symbols) equals the viscous one (hollow symbols) in a broad interval of frequencies (see Figure 3), criterion also established by Tung [21] to determine the gel point. In this figure, the time dependence of the elastic and viscous moduli is plotted for 3 different frequencies: 1 rad/s (black circles), 10 rad/s (red squares), and 100 rad/s (green triangles), and we can see here that the time when $G' = G''$ is placed between 31 and 32 hours (vertical lines in Figure 3). As we can see the values of the gel time obtained by both methods coincide, showing the validity of our analysis. Moreover, the critical viscoelastic exponents found in our system is in accordance with the interval 0.6–0.8 reported by other authors for this critical exponent in tetraethyl orthosilicate based hybrid solutions [32, 33].

To deepen the study of the gel point, we proceed with the technique previously described, known as microrheology. Passive microrheology is a technique that allows determination of the viscoelastic properties of materials with no need of applying external forces to the sample, since the only forces exerted on the system are those due to the thermal motion of the tracer particles immersed in the medium. In this way, the deformation energy in the system is of the order of the thermal energy that is responsible for the Brownian motion

of the probe particles, guaranteeing the linear viscoelastic regime. Sometimes, the inconvenience of this technique is either the difficulty of finding suitable tracers since they must not alter the physical properties of the medium that will be characterized or the mechanical agitation occasionally needed to disperse the particles. In our case, we can circumvent these problems since there is no need to introduce the tracer particles to the system, by using as tracers the ZnO particles that are already present in our material. In microrheology experiments based on light scattering, the autocorrelation function of the intensity scattered by the tracers is measured. Here it is critical to ensure that most of the light scattered by the sample comes from the optical probes, as it happens in our system due to the ZnO particle concentration used in the sample preparation, and the low scattering power of the suspending medium. In this case, the autocorrelation function only contains information on the tracer motion. For example, when the probes particles are immersed in a Newtonian fluid, the autocorrelation function of the intensity minus one decays exponentially and its decay time is proportional to the medium viscosity; see (2) and (4). In our case, we use this relation to follow the time evolution of the viscosity of our polymeric material by measuring the autocorrelation function of the intensity scattered by the ZnO particles immersed in the dispersion of interest, as a function of reaction time, and data are shown in Figure 4(a). Here we can see that for the initial 30 hours the intensity autocorrelation function decays exponentially, meaning that the suspending medium is mostly viscous, and the only difference that can be appreciated is the increase of the decay time as the system evolves. This increase of the decay time, since the size particle is kept constant along the experiment, could only be due to an increase of the medium viscosity. However, for waiting times longer than 31 hours, we can appreciate that the autocorrelation function does not decay to zero. This feature is due to the fact that the ZnO tracer particles immersed in the polymer dispersion are dynamically arrested; thus they cannot move distances that allow light decorrelation. The dynamic arrest of the tracers, inferred from these measurements, is a clear evidence of the emergence of a three-dimensional network in the suspending medium. Quantitatively, this liquid to solid transition can be analyzed by plotting the decay time of $g_I(\tau) - 1$ normalized by the decay time of the autocorrelation function measured just after sample preparation as a function of waiting time for the ergodic samples, as plotted in Figure 4(b). The ratio between these two decay times corresponds precisely to the system viscosity normalized by the viscosity of the sample at time zero. Here we can notice that the time behavior of the sample viscosity obtained by light scattering resembles that measured by rotational rheology, which was plotted in Figure 1(b). In fact, the experimental time dependence of the viscosity shown in the figure can be very well described by the same power law, (5). From the fit, we obtain that the gelation time is $t_{\text{gel}} = (30.1 \pm 0.2)$ hours, and $s = (0.62 \pm 0.03)$, values that are very similar to those found from rotational rheometry.

This gel time value is smaller than the one obtained for the polymeric matrix without nanoparticles, which is around 35 hours. Therefore, as found for CuO nanoparticles in this type

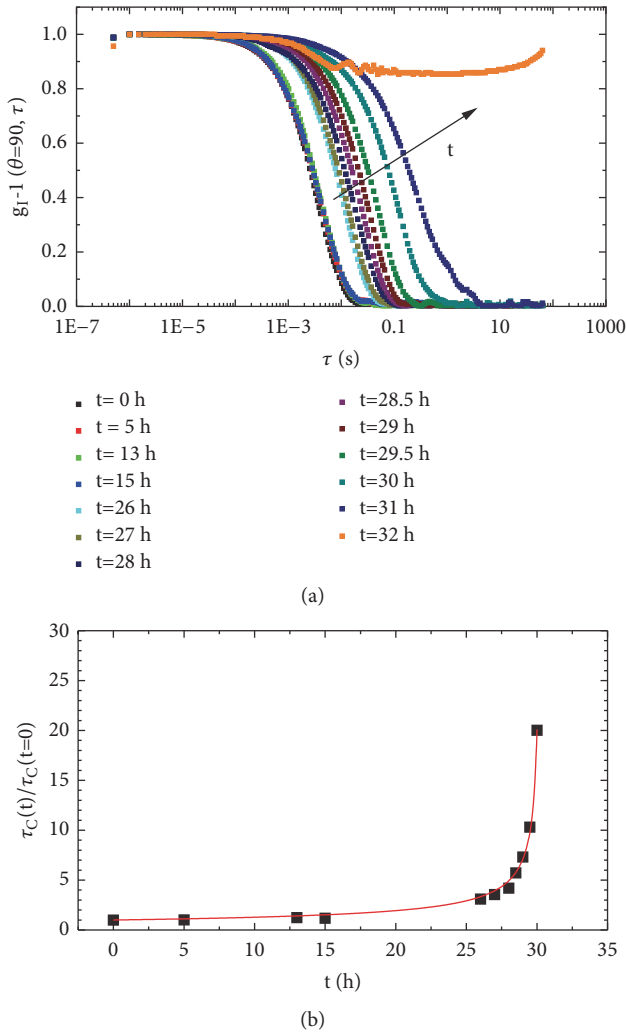


FIGURE 4: (a) Intensity autocorrelation function minus one measured as a function of time. (b) Ratio of the characteristic times obtained from the analysis of the autocorrelation functions shown in (a) normalized by the initial decay time measured at $t = 0$ h. Line is a fit to a critical power law: $\tau_C(t)/\tau_C(t=0) = (1 - (t/t_{gel}))^{-s}$, where τ_0 is the decay time of the autocorrelation at time $t = 0$ h and t_{gel} and s are fitting parameters.

of formulations [17], we can think this acceleration is due to the favoured release of Zn^{2+} ions at the basic conditions of the initial sol, which accelerates the hydrolysis of the ethoxy groups, and to the fact that Zn^{2+} may also interact with $[Si(OH)_X]^-$ intermediate species, which promotes the formation of SiO_2 nuclei by condensation.

4. Conclusions

In this work, we present a study on the gelation process of an organically modified silicate that consists of a silica polymeric precursor mixed with n-octylamine and an organosilicon compound, PDMS, which has application in restoration and conservation of building substrates. During material preparation, ZnO particles are introduced into the system in order

to obtain stone treatments that, besides their consolidant and hydrofugant properties, present biocide activity [12, 13]. Moreover, the resulting product shows low viscosity that ensures an easy and homogenous application over the stone and a reasonable reaction time to reach the gel point. Despite the importance of the gel point in material science, its determination is still controversial since the sol-gel transition can be found at different points depending on the method employed. For this reason, we have determined the gel point of our sol-gel process by different techniques: rotational and oscillatory rheology and light scattering experiments. The gel point estimated from the divergence of the viscosity, measured by rotational rheology, is $t_{gel} = (30.2 \pm 0.1)$ hours, and it is very similar to that obtained from the time where the frequency dependence of the viscoelastic moduli obeys a power law $t_{gel} = 31$ hours and from the crossing point where the elastic modulus equals the viscous one, between 31 and 32 hours. The critical viscoelastic exponent found from the frequency dependence of the storage and loss moduli at the gel point is around (0.81 ± 0.07) and (0.73 ± 0.02) , respectively, values similar to the expected value of 0.67 obtained from percolation theory. Furthermore, by performing light scattering experiments where the sample is not perturbed, we obtain $t_{gel} = (30.1 \pm 0.2)$ hours that agrees with those found from rheological measurement. Here we show how light scattering, a noninvasive technique, can be applied to study a new family of gelling products that contain nanoparticles for a given application by using them as tracers to characterize the mechanical properties of such products. Although biocide activity of the composite presented here has not yet been proved, this work is an attempt to obtain stone coatings with biocide activity that do not alter the stone appearance and still have consolidating and hydrofugant properties. In this direction, further studies using different ZnO particle concentrations and different size of particles are currently under development.

Data Availability

The data used to support the findings of this study are available from the corresponding author upon request.

Conflicts of Interest

The authors declare that there are no conflicts of interest regarding the publication of this paper.

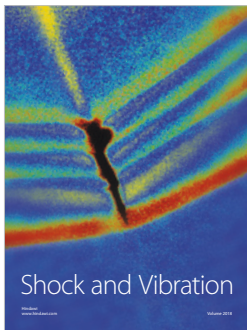
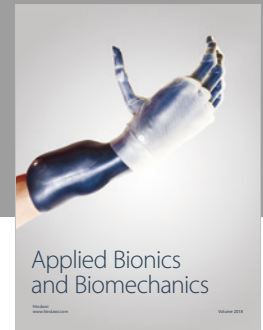
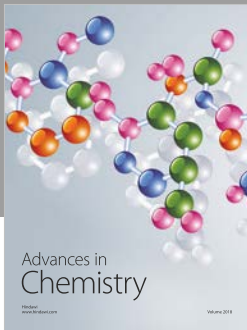
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